

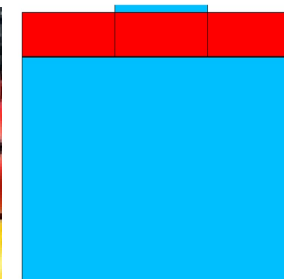
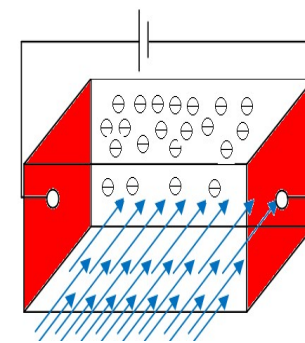
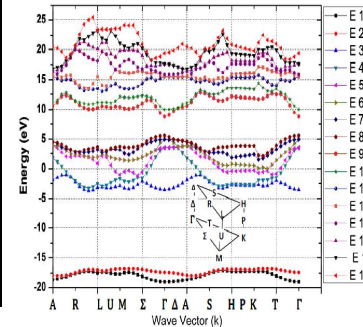
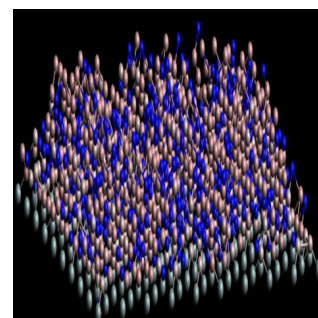


MOCVD Vertical Flow Reactor

TNL-Showerhead Simulator



*Technology of Next Level
driven through innovation*



SUBSTRATES AVAILABILITY



SILICON

- Available in up to >30 cm diameter
- Quite inexpensive and high quality
- Can be obtained n -type, p -type, or with high resistivity
- Used for Si and SiGe technologies
- Intense research to develop Si-based "pseudo-substrates" for GaAs, InP, CdTe...technologies

GaAs

- Available in up to >12 cm diameter
- High quality, more expensive than Si, but affordable
- Used for GaAs and AlGaAs, and strained InGaAs technologies
- Can be used for electronic and optoelectronic applications

InP

- 10 cm diameter available, but expensive
- InP and InGaAsP technologies can be grown
- Very important for optoelectronics and high performance electronics

SiC

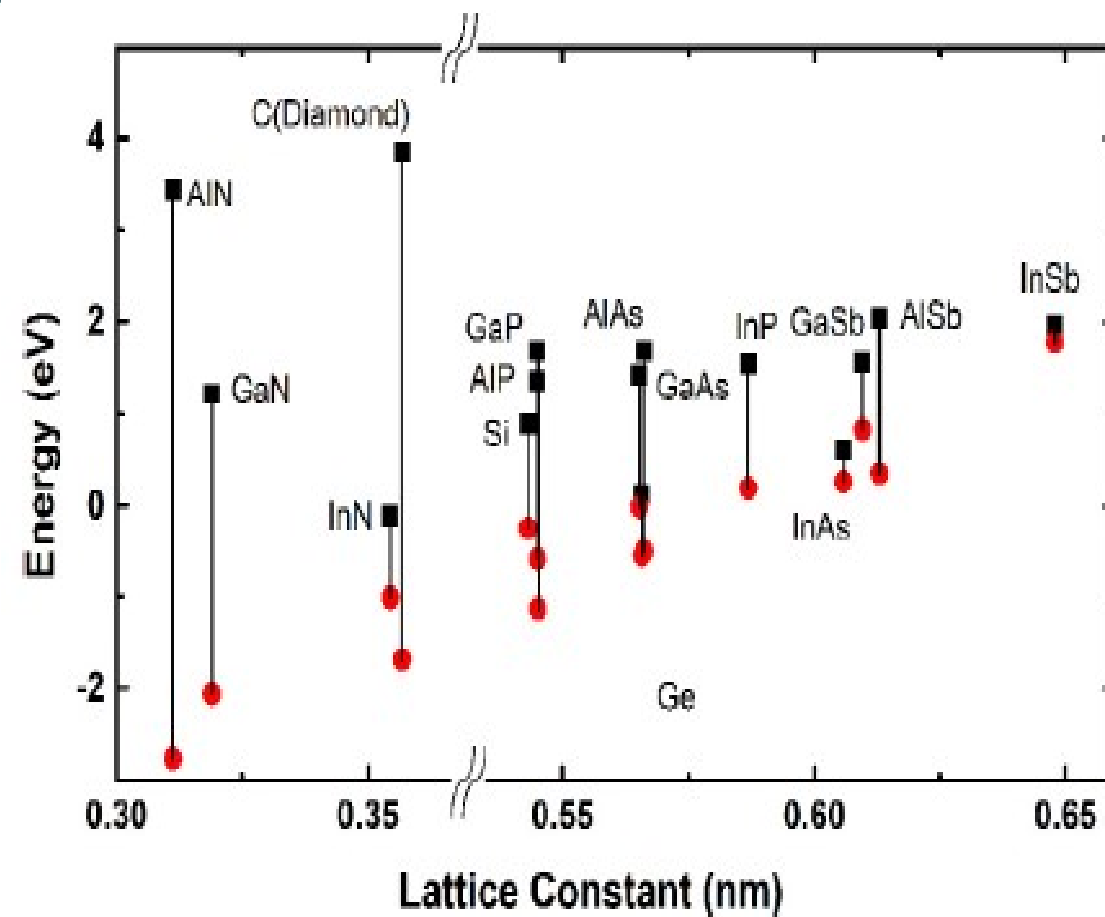
- Small, very expensive substrates
- Very important for high power, large gap technologies
- Used for nitride technology

1. Bulk Crystal Growth

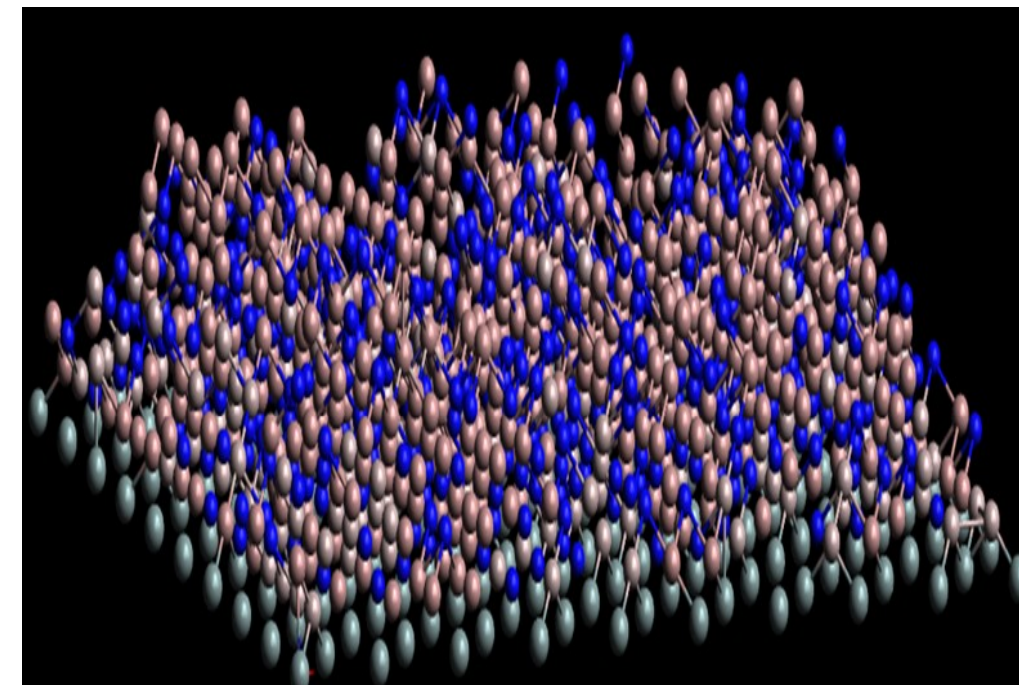
- ▶ State of the art device technologies depends on: **Purity & Perfection of the crystals**
- ▶ Limited to **Si**, **GaAs** and upto some extent for **InP**



EPITAXIAL GROWTH CHALLENGES



Development of pseudo-GaAs substrates over Si



*E.T. Yu, J.O. McCaldin, T.C. McGill, Band offsets in semiconductor heterojunctions, in: E. Henry, T. David (Eds.), Solid State Physics, Academic Press, 1992, pp. 1-146

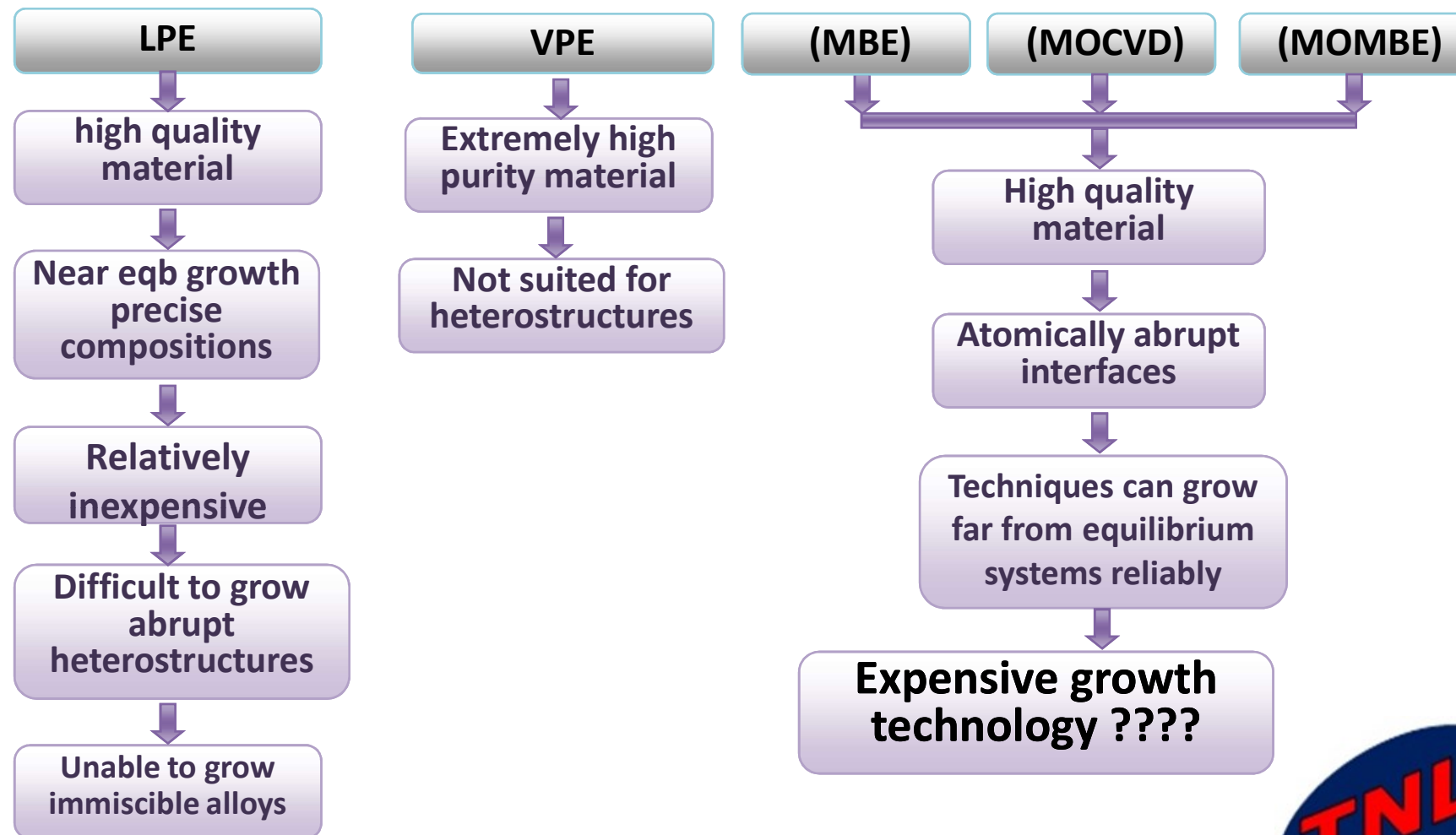


AVAILABLE TECHNIQUES



Epitaxial Growth Challenges

- Semiconductor technologies dependent on non ideal substrates
- Lot of Technological Challenges and issues
- Need for Pseudo-substrates
- Low Defect densities
- Strain Engineering



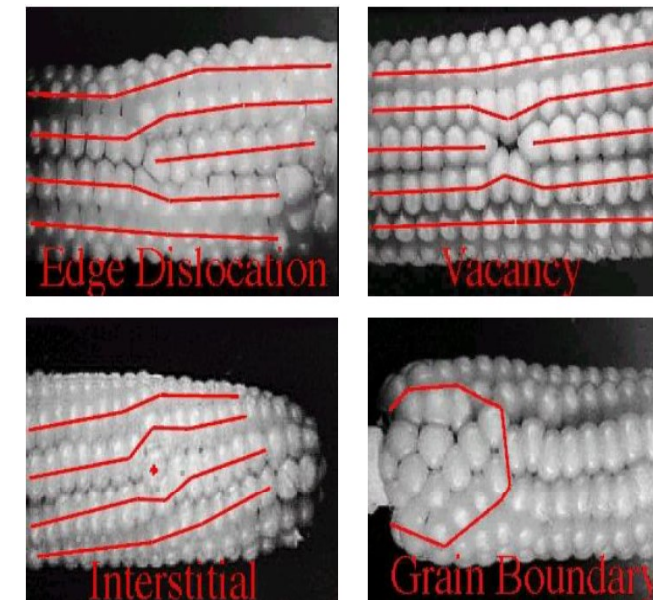
CHALLENGES



E.g. GaN growth over various substrates:

Substrate	Si	Al ₂ O ₃	SiC	Bulk GaN	AlN
Lattice Mismatch (%)	17	16	3.4	-	2.5
Thermal Conductivity (W/mm-k)	150	35	490	260	319
Resistivity (ohm-cm)	10 ⁴	10 ¹⁴	~10 ¹²	-	>10 ¹⁴

Defects



Chungnam Natl Univ., S. K. Hong

- Group IV, III-V, II-VI epi-growth with multi components
- Point defects, e.g. vacancies, interstitial atoms
- Extended defects within the film, generally dislocations and stacking faults
- Dislocations: reduce or relax strain through lattice mismatch or thermal expansion differences.

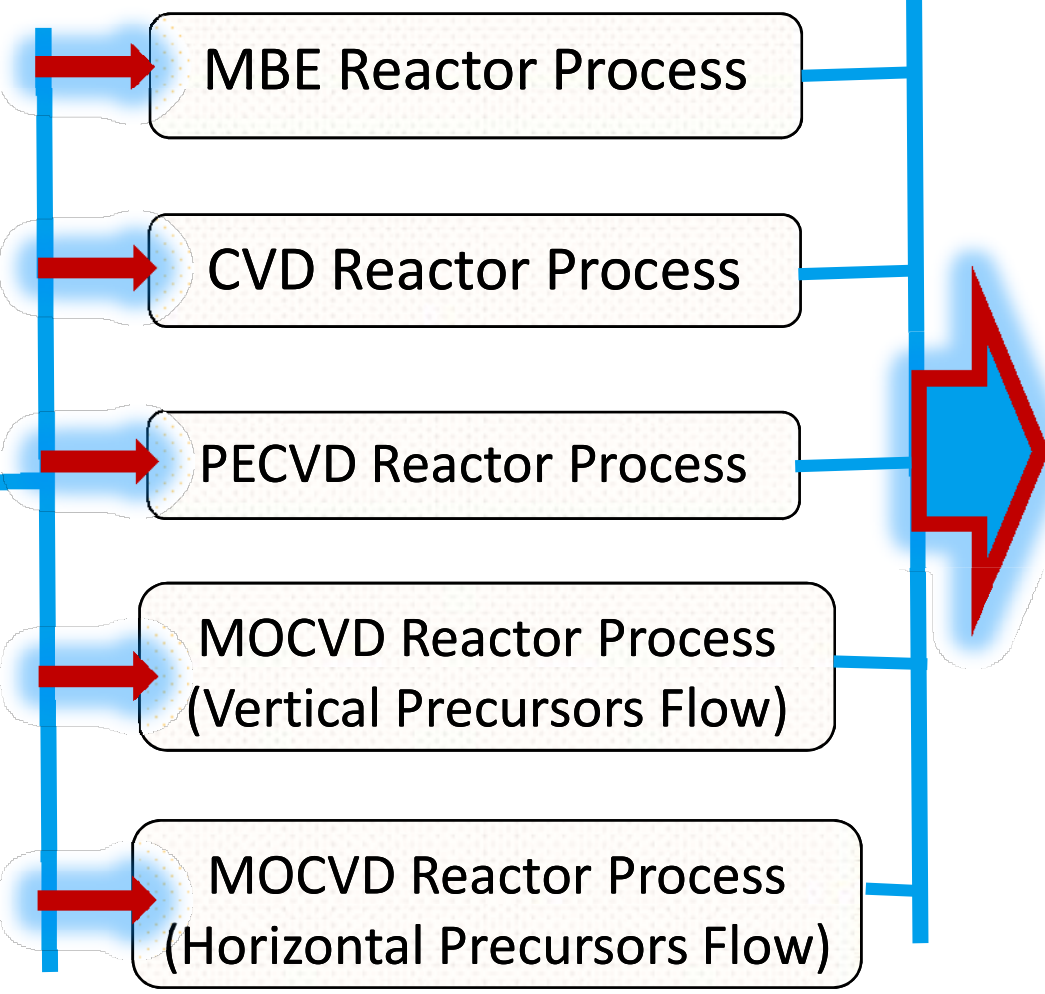


EPITAXY PROCESS SOLUTION



Innovative Atomistic Scale Reactor Simulation without use of Continuum models

TNL Framework



- User's Input Growth Conditions
- Surface Profile (Roughness)
- Strain Mapping (layer by Layer)
- Lattice parameter
- Defects (Vacancies, Interstitials, Dislocations, Stacking Faults)



INPUTS : MOCVD PROCESS



Chamber Condition

Showerhead Based

Injector Based

Viscosity (poise)

Mass Diffusivity (cm²/s)

Chamber Radius (cm)

ShowerHead Parameters

Shower hole's diameter

Chamber Volume (ltrs.)

Chamber Pressure torr

Ceiling Height (cm)

Chamber Temperature (C)

Sticking Coeff.

Precursor Condition

Number of Port

Precursor 1

Flow Rate atm cc/s

Step 1

Precursor Condition

Number of Port

Precursor 1

Flow Rate atm cc/s

- Select Precursor
- Ga(CH₃)₃
- (CH₃)₃CA₅H₂
- H₂
- Al(CH₃)₃
- NH₃
- CH₃
- O₂

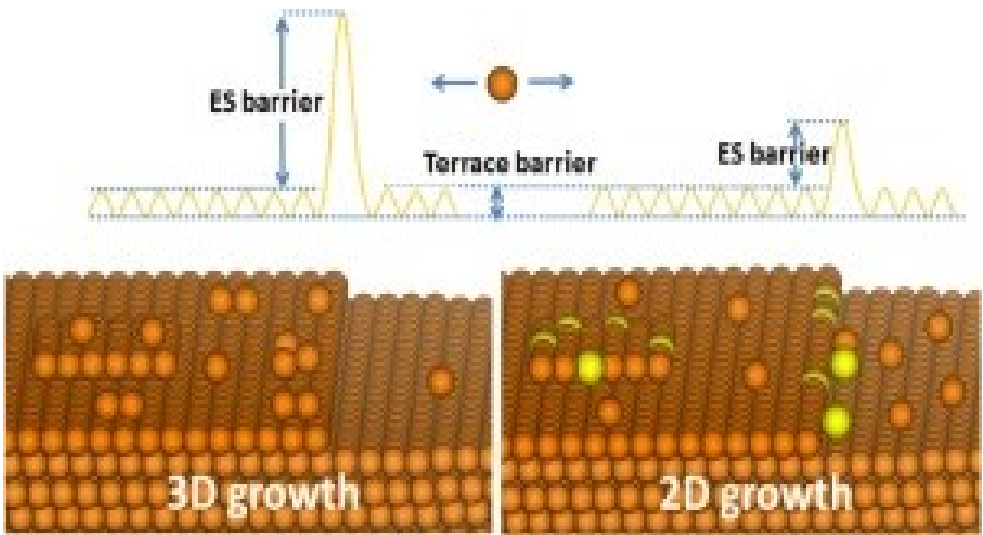
Many More parameters details Require



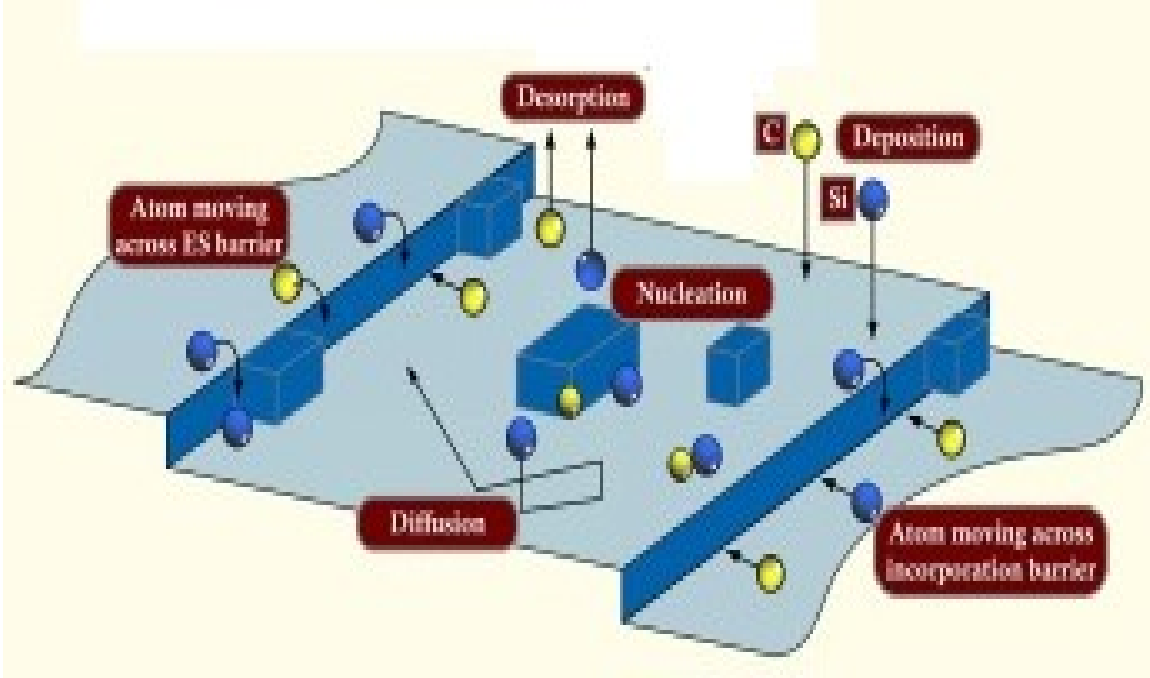
METHODOLOGY: MOCVD PROCESS



Schwoebel barrier: The atom diffuses from the site exactly above the edge atom to the site immediately next to the edge atom as;



Incorporation barrier: The atom incorporates into the edge on the same surface level.



CHEMICAL KINETICS SOLUTION



- TNL Chemical Kinetics database includes gas- and surface phase chemical reactions
- Users may choose any desired equation or set of equations for the precursors they input based on requirements
- Users have flexibilities to write their own chemical reactions



TNL_Chemical Kinetics

Precursors

H2,SiH4

No.	Name	A	n	E(Cal)
1	G 1 SiH4 --> SiH2 + H2	9.49	1.7	54710
2	G 2 SiH4 + SiH2 --> Si2H6	10.26	1.7	50200
3	G 3 Si2H6 + SiH2 --> HSiSiH3 + ...	14.24	0.4	8900
4	G 4 Si2H6 --> H2 + HSiSiH3	9.96	1.8	54200
5	G 5 HSiSiH3 --> H2SiSiH2	13.40	0.2	5380
6	G 6 HSiSiH3+H2 --> SiH2 + SiH4	13.97	0	4092

No.	Name	A	n	E(Cal)
2	S 2 SiH2 + sigma --> Si+H2	11.76	0.5	0
3	S 3 H2 + 2sigma --> 2H*	11.36	0.5	17250

Output_Window

No.	Gas_Reaction	A	n	E(Cal)
1	G 1 SiH4 --> SiH2 + H2	9.49	1.7	54710
2	G 2 SiH4 + SiH2 --> Si2H6	10.26	1.7	50200
3	G 3 Si2H6 + SiH2 --> HSiSiH3 + SiH4	14.24	0.4	8900
4	G 4 Si2H6 --> H2 + HSiSiH3	9.96	1.8	54200
5	G 5 HSiSiH3 --> H2SiSiH2	13.40	0.2	5380

No.	Surface_Reaction	A	n	E(Cal)
2	S 2 SiH2 + sigma --> Si+H2	11.76	0.5	0
3	S 3 H2 + 2sigma --> 2H*	11.36	0.5	17250

Editor

TMG, NH₃ & N₂/H₂/ARGON



Gas-phase Mechanisms Reactions

	$k = AT^n e^{-E_a/RT}$					A	n	E_a
G1	TMG	=	DMG	+	CH ₃	1.00×10^{47}	-9.18	76,996
G2	DMG	=	MMG	+	CH ₃	7.67×10^{43}	-9.8	34,017
G3	MMG	=	Ga	+	CH ₃	1.68×10^{30}	-5.07	84,030
G4	TMG	+	NH ₃	→	TMG:NH ₃	2.28×10^{34}	-8.31	3115
G5	TMG	+	NH ₃	→	DMG:NH ₂	1.70×10^4	2	19,969
G6	DMG	+	NH ₃	→	DMG:NH ₃	4.08×10^{31}	-7.03	3234
G7	DMG	+	NH ₃	→	MMG:NH ₂	5.30×10^5	1.56	20,744
G8	MMG	+	NH ₃	→	MMG:NH ₃	7.95×10^{24}	-5.21	2094
G9	MMG	+	NH ₃	→	GaNH ₂	8.10×10^5	1.3	17,722
G10	NH ₃	+	CH ₃	→	NH ₂	3.31×10^3	2.51	9859
G11	CH ₃	+	H ₂	→	CH ₄	1.20×10^{12}	0	12,518
G12	TMG	+	H	→	DMG	5.00×10^{13}	0	10,036
G13	DMG	+	H	→	MMG	5.00×10^{13}	0	10,036
G14	TMG:NH ₃	→	MMG	+	2CH ₃	1.33×10^{44}	-8.24	77,791
G15	CH ₃	+	H	+	M	2.40×10^{22}	-1	0
G16	2CH ₃	=	C ₂ H ₆			2.00×10^{13}	0	0
G17	2H	+	M	=	H ₂	2.00×10^{16}	0	0



TMG, NH₃ & N₂/H₂/ARGON



Surface phase Reactions: PATH 1

Path 1, $k = AT^n e^{-E_a/RT}$					A	n	E _a	
1	MMG	+	N(S)	→	MMG(S)	1.16×10^5	2.98	0
2	MMG(S)	→	MMG	+	N(S)	1.12×10^{14}	0.55	107,673
3	NH ₃	+	MMG(S)	→	COMPM1(S)	3.35×10^7	3.33	0
4	COMPM1(S)	→	NH ₃	+	MMG(S)	5.70×10^{13}	-0.16	8146
5	MMG	+	COMPM1(S)	→	CH ₄ + COMPM2(S)	1.23×10^{10}	3.22	23,446
6	NH ₃	+	COMPM2(S)	→	COMPM3(S)	3.35×10^7	3.33	0
7	COMPM3(S)	→	NH ₃	+	COMPM2(S)	5.70×10^{13}	-0.161	8146
8	MMG	+	COMPM3(S)	→	CH ₄ + COMPM4(S)	1.23×10^{10}	3.22	23,446
9	NH ₃	+	COMPM4(S)	→	COMPM5(S)	3.35×10^7	3.33	0
10	COMPM5(S)	→	NH ₃	+	COMPM4(S)	5.70×10^{13}	-0.161	8146
11	COMPM5(S)	→	CH ₄	+	RINGM1(S)	1.23×10^7	3.22	23,446
12	Ga(S)	+	RINGM1(S)	→	RINGM2(S) + N(S)	3.35×10^7	3.33	0
13	RINGM2(S)	→	3H ₂	+	3GaN(B) + Ga(S)	3.68×10^9	2.05	59,610



TMG, NH₃ & N₂/H₂/ARGON



Surface phase Reactions: PATH 2

	Path 2, $k = AT^n e^{-E_a/RT}$				A	n	E _a	
14	CH ₃	+	Ga(S)	→	MMG(S)	1.76×10^9	1.39	0
15	MMG(S)	→	CH ₃	+	Ga(S)	4.54×10^{13}	0.0346	79,480
16	NH ₂	+	Ga(S)	→	NH ₂ (S)	3.17×10^8	1.83	0
17	GaNH ₂	+	N(S)	→	GaNH ₂ (s)	2.27×10^6	2.247	0
18	GaNH ₂ (S)	→	GaNH ₂	+	N(S)	4.83×10^{13}	0.614	83,881
19	COMPMM1(S)	→	CH ₄	+	GaNH ₂ (S)	1.49×10^{11}	0.609	25,950
20	MMG	+	GaNH ₂ (S)	→	COMPMM1(S)	1.16×10^5	2.98	0
21	NH ₃	+	COMPMM1(S)	→	COMPMM2(S)	3.35×10^7	3.33	0
22	COMPMM2(S)	→	CH ₄	+	COMPMM3(S)	1.49×10^{11}	0.609	25,950
23	MMG	+	COMPMM3(S)	→	COMPMM4(S)	1.16×10^5	2.98	0
24	NH ₃	+	COMPMM4(S)	→	COMPMM5(S)	3.35×10^7	3.33	0
25	COMPMM5(S)	→	CH ₄	+	RINGM1(S)	1.49×10^{11}	0.609	25,950
26	NH ₂ (S)	→	NH ₂	+	Ga(S)	1.45×10^{14}	0.09	59,786
27	COMPMM1(S)	→	MMG	+	GaNH ₂ (S)	1.00×10^{14}	0.55	42,819
28	COMPMM2(S)	→	NH ₃	+	COMPMM1(S)	5.70×10^{13}	-0.1	8146
29	COMPMM4(S)	→	MMG	+	COMPMM3(S)	1.00×10^{14}	0.55	42,819
30	COMPMM5(S)	→	NH ₃	+	COMPMM4(S)	5.70×10^{13}	-0.1	8146
31	Ga	+	N(S)	→	Ga(S)	1.00×10^{11}	1.5	0
32	Ga(S)	+	NH ₂ (S)	→	GaNH ₂ + Ga(S)	1.00×10^{25}	0	0
33	Ga(S)	→	Ga	+	N(S)	1.00×10^{13}	0	45,168
34	6CH ₃	+	RINGM2(S)	→	COM1(S)	7.55×10^7	2.31	0
35	COM1(S)	→	6CH ₃	+	RINGM2(S)	1.00×10^{13}	0.71	45,506
36	COM1(S)	→	6CH ₄	+	3GaN(B) + Ga(S)	4.00×10^{12}	0	49,675



TMG, NH₃ & N₂/H₂/ARGON



Surface phase Reactions: PATH 3

	Path 3, $k = AT^n e^{-E_a/RT}$					A	n	E _a
37	TMG	+	N(S)	→	TMG(S)	1.16×10^5	2.98	0
38	NH ₃	+	TMG(S)	→	TCOM1(S)	3.35×10^7	3.33	0
39	TCOM1(S)	→	CH ₄	+	TCOM2(S)	1.49×10^{11}	0.609	32,785
40	Ga(S)	+	TCOM2(S)	→	TCOM3(S)	3.35×10^7	3.33	0
41	TCOM3(S)	→	2CH ₄	+	GaN(B)	1.49×10^{11}	0.609	49,675
42	TMG(S)	→	TMG	+	N(S)	1.12×10^{14}	0.55	49,675
43	TCOM1(S)	→	NH ₃	+	TMG(S)	5.70×10^{13}	-0.161	11,922
44	TMG:NH ₃	+	N(S)	→	TCOM1(S)	1.16×10^5	2.98	0
45	TCOM1(S)	→	TMG:NH ₃	+	N(S)	1.12×10^{14}	0.55	49,675
46	TCOM1(S)	→	2CH ₃	+	MMG(S)	1.12×10^{14}	0.55	10,7673
47	MMGNH ₃	+	N(S)	→	COMPM1(S)	1.16×10^5	2.98	0
48	COMPM1(S)	→	MMG:NH ₃	+	N(S)	1.12×10^{14}	0.55	107,673
49	MMG:NH ₃	+	COMPM1(S)	→	CH ₄	1.23×10^{10}	3.22	23,446
50	MMG:NH ₃	+	COMPM3(S)	→	CH ₄	1.23×10^{10}	3.22	23,446
51	MMG:NH ₃	+	GaNH ₂ (S)	→	COMPMM2(S)	1.16×10^5	2.98	0
52	MMG:NH ₃	+	COMPMM3(S)	→	COMPMM5(S)	1.16×10^5	2.98	0



COMPOSITION OF INTERMEDIATES



Chemical Composition of compound on the surface

Compounds Names	Chemical Formula
COMPM1(S)	$\text{NH}_3 \cdot \text{MMG(S)}$
COMPM2(S)	$\text{Ga} \cdot \text{NH}_2 \cdot \text{MMG(S)}$
COMPM3(S)	$\text{NH}_3 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{MMG(S)}$
COMPM4(S)	$\text{Ga} \cdot \text{NH}_2 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{MMG(S)}$
COMPM5(S)	$\text{NH}_3 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{MMG(S)}$
RINGM1(S)	$\text{NH}_2 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{Ga(S)}$
RINGM2(S)	$(\text{S})\text{NH}_2 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{Ga(S)}$
COMPMM1(S)	$\text{MMG} \cdot \text{GaNH}_2(\text{S})$
COMPMM2(S)	$\text{NH}_3 \cdot \text{MMG} \cdot \text{GaNH}_2 \cdot \text{Ga(S)}$
COMPMM3(S)	$\text{NH}_2 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{Ga(S)}$
COMPMM4(S)	$\text{MMG} \cdot \text{NH}_2 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{Ga(S)}$
COMPMM5(S)	$\text{NH}_3 \cdot \text{MMG} \cdot \text{NH}_2 \cdot \text{Ga} \cdot \text{NH}_2 \cdot \text{Ga(S)}$
TCOM1(S)	$\text{NH}_3 \cdot \text{TMG(S)}$
TCOM2(S)	$\text{NH}_2 \cdot \text{DMG(S)}$
TCOM3(3)	$(\text{S})\text{NH}_2 \cdot \text{DMG(S)}$
COM1(S)	$\text{RINGM2(S)} \cdot \text{CH}_3 \text{ complex}$



CASE STUDY: GaAs OVER GaAs *

*P. K. Saxena, P. Srivastava, R. Trigunayat, An innovative approach for controlled epitaxial growth of GaAs in real MOCVD reactor environment, [*Journal of Alloys and Compounds*](#) 809 (2019) 151752.

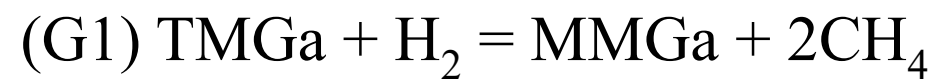


CHEMICAL KINETICS



Reactions included in the gas-phase
and surface-phase model

Gas-phase reactions:



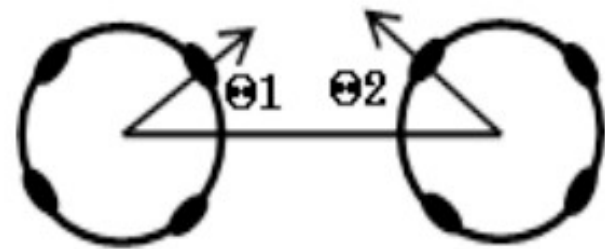
Surface reaction:



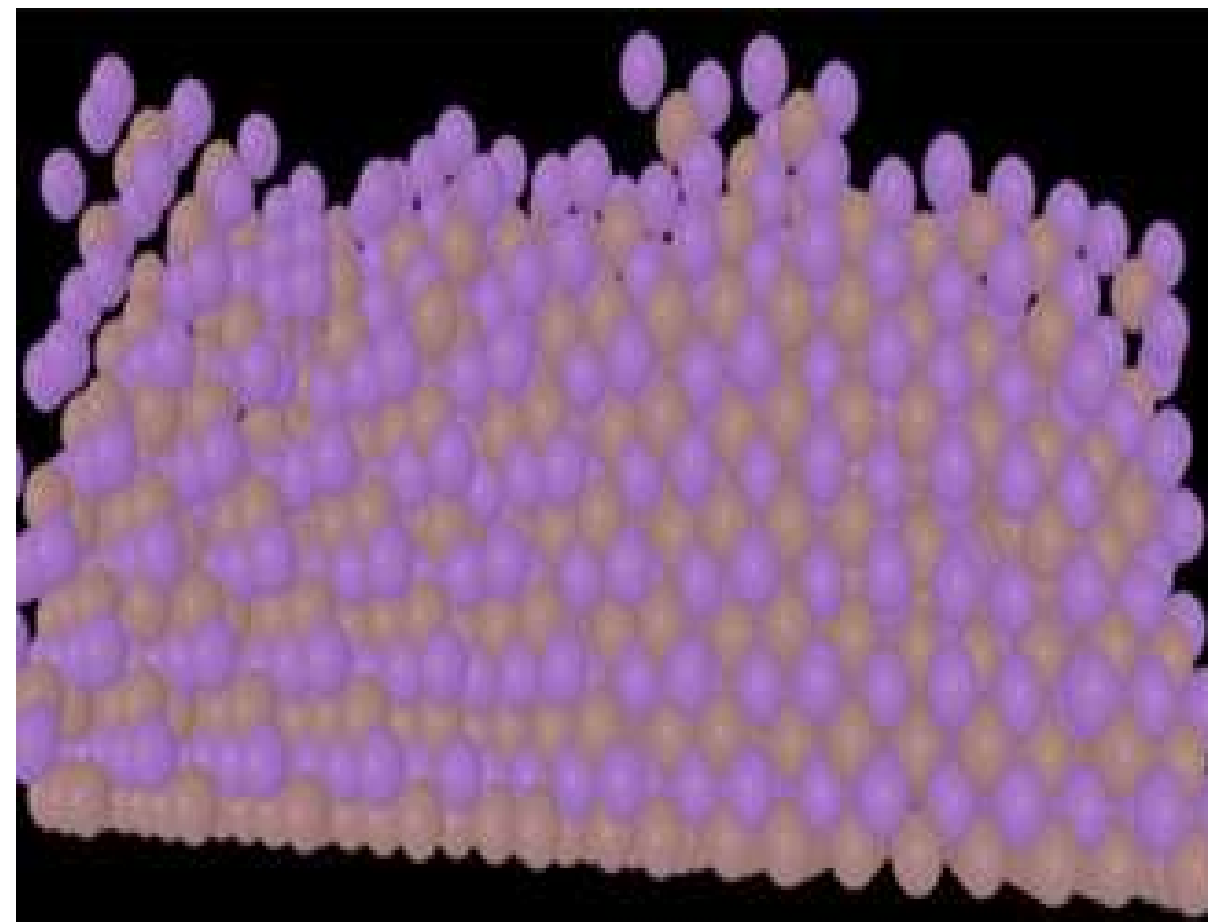
Kinetic Parameter	Value	Unit
A_{G1}	1.2×10^{15}	S^{-1}
A_{G2}	5.32×10^{15}	S^{-1}
A_{S1}	1.23×10^9	m/s
E_{G1}	196	kJ/mol
E_{G2}	203	kJ/mol
E_{S1}	130	kJ/mol



MOCVD GAAS OVER GAAS



An atom makes bond with another atom when its patches lie in a line. The interaction between two atoms depends upon the angle its patches make with each other.

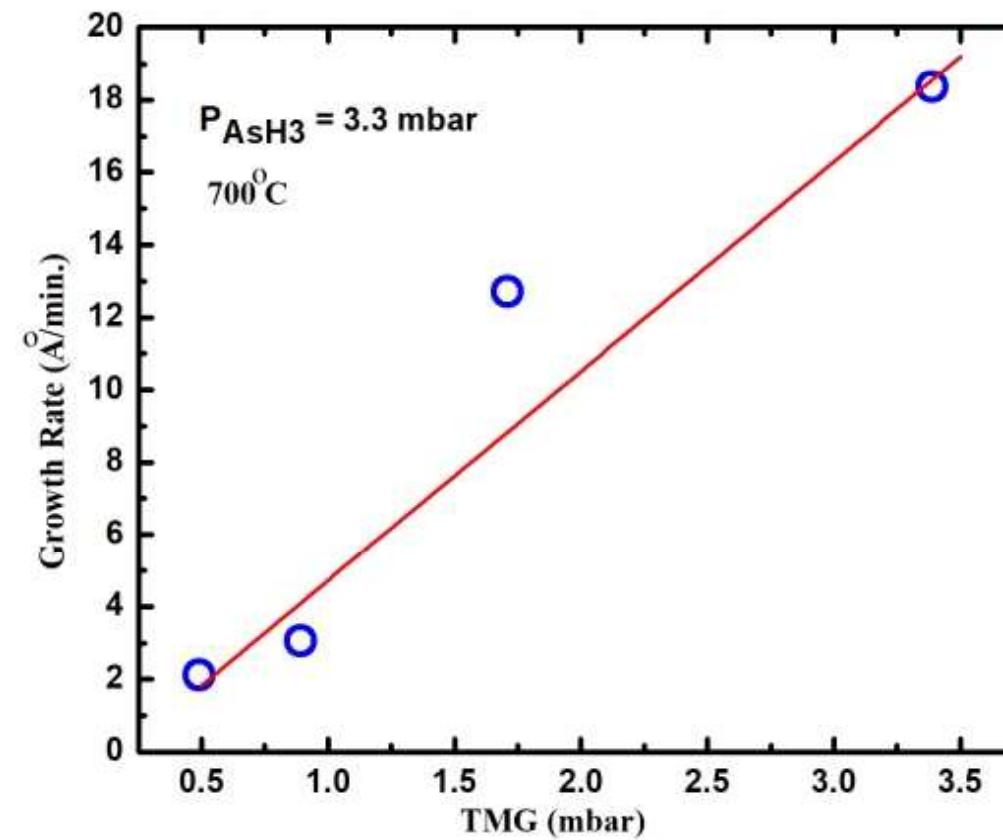


* [Journal of Alloys and Compounds 809 \(2019\) 151752.](#)

MOCVD Growth Rate



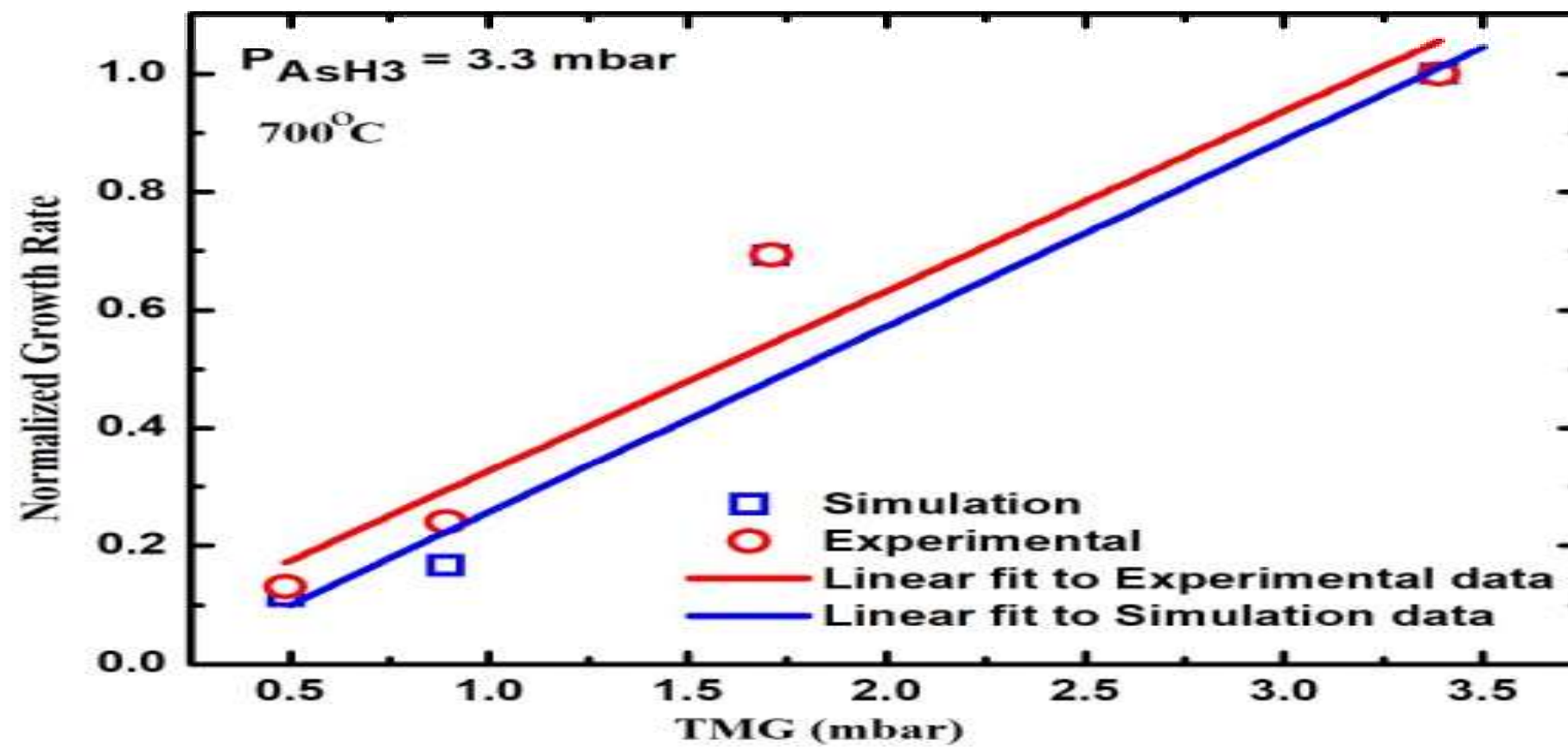
Parameters	Values	Units
Substrate surface energy (E_s)	1.75	eV
Schwoebel barrier energy (E_{shw})	0.02	eV
Incorporation barrier energy (E_i)	0.05	eV



* [Journal of Alloys and Compounds 809 \(2019\) 151752.](#)



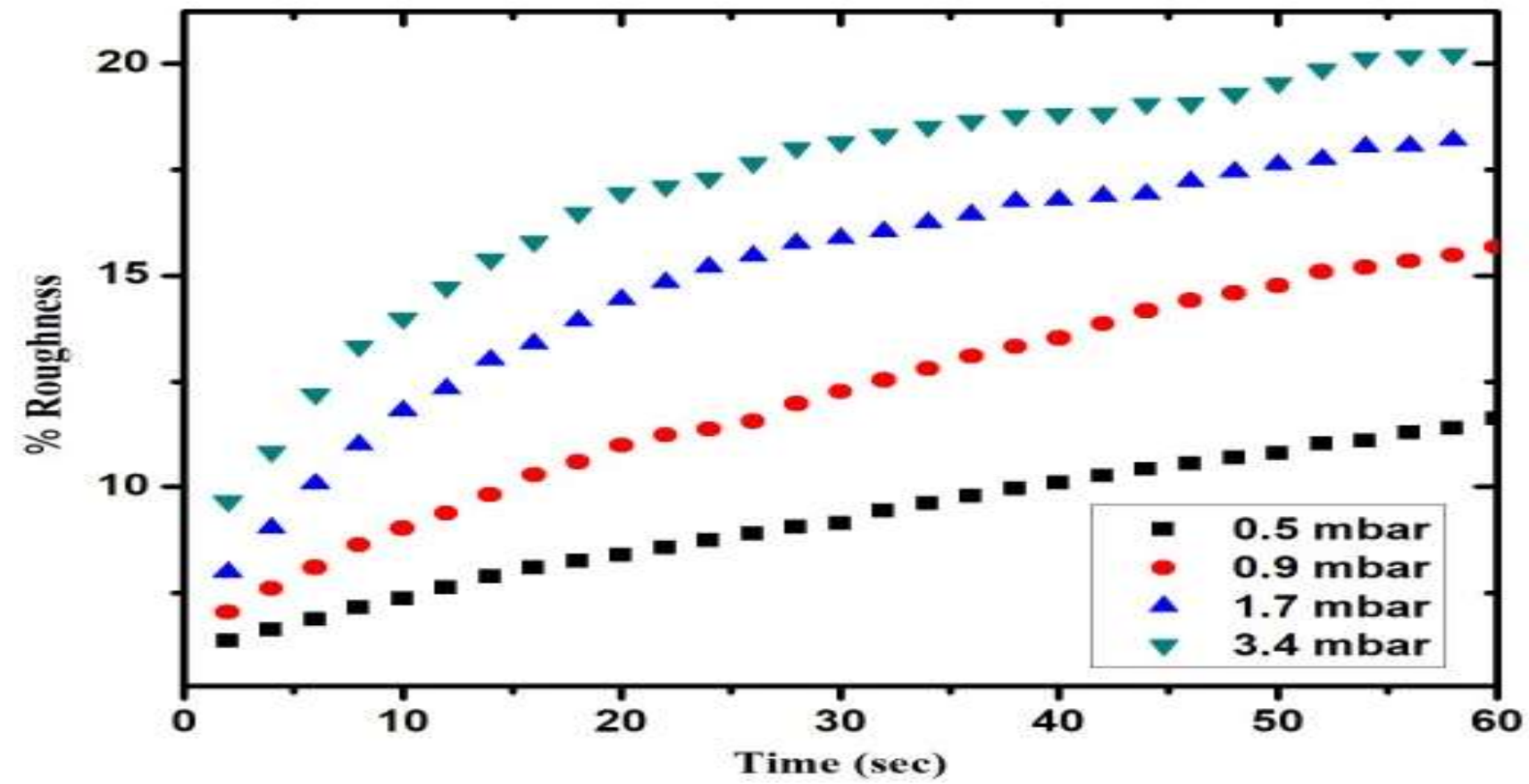
MOCVD Growth Rate



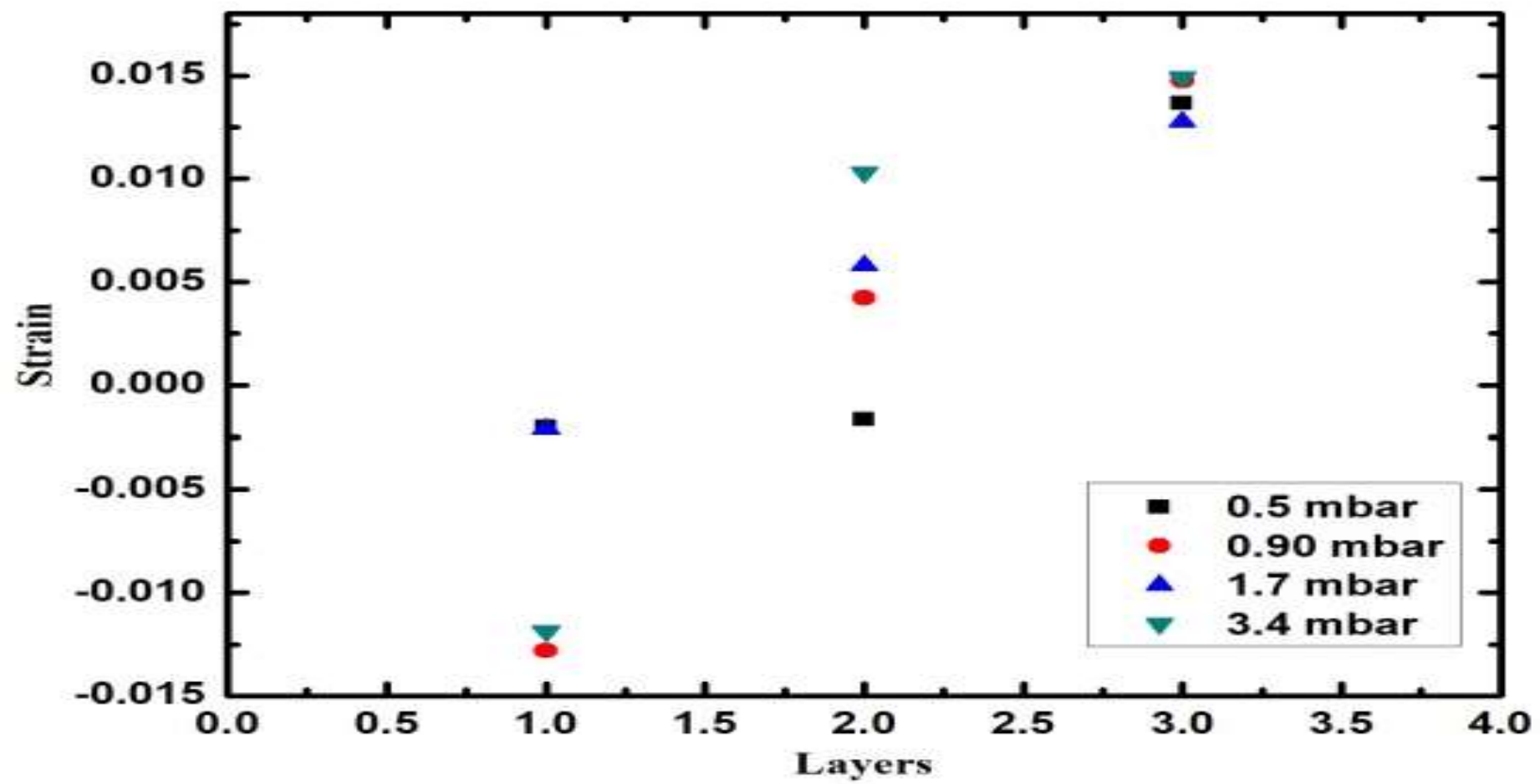
* [Journal of Alloys and Compounds 809 \(2019\) 151752.](#)



ROUGHNESS



STRAIN



ADVANCE LICENSING & PRICE VALUE



TNL's tools support advanced and unique licensing models tailored for unique customer needs.

➤ **ADVANCED LICENSING OPTIONS:**

- Term-Based
- Perpetual
- TCAD Academic Suite
- 24x7 Technical Support for **Academic Institutions**



Thank You
Contact us



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